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The Crystal Structure of a New Ferroelectric Compound, $\text{NaTh}_2(\text{PO}_4)_3$ **

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The crystal structure of ferroelectric $\text{NaTh}_2(\text{PO}_4)_3$ has been determined using X-ray diffraction techniques. The crystals are monoclinic with dimensions $a = 17.37(2)$, $b = 6.81(1)$, $c = 8.13(1) \text{ \AA}$, $\beta = 101.03^\circ$ and belong to the space group Cc. The density calculated for $Z = 4$ is $5.435 \text{ g} \cdot \text{cm}^{-3}$; that determined by a picnometer is $5.41 \text{ g} \cdot \text{cm}^{-3}$. The structure was solved from 913 observed, independent reflections. The atomic coordinates and temperature factors were refined by full-matrix least squares. The discrepancy indices had final values of 7.1% (including unobserved reflections) and 6.2% (omitting unobserved reflections). The compound may be designated as sodium trisphosphatodithorate. The thorium atom is nine coordinated and thorium-oxygen distances are from 2.27 to 2.58 Å, and phosphorus-oxygen distances vary from 1.52 to 1.59 Å. The atomic arrangement is isostructural with that of $\text{KTh}_2(\text{PO}_4)_3$. The main difference in the structures of $\text{NaTh}_2(\text{PO}_4)_3$ and $\text{KTh}_2(\text{PO}_4)_3$ is the position of the alkali metal atom. The potassium atom in $\text{KTh}_2(\text{PO}_4)_3$ is on the special position 0, y , $1/4$ on the twofold axis in the space group C2/c, and the sodium atom in $\text{NaTh}_2(\text{PO}_4)_3$ is out of twofold axis for 0.8 Å.

Introduction

X-ray diffraction data,¹ the preparation of single crystals and crystallographic data of $\text{NaTh}_2(\text{PO}_4)_3$ ² have been published already. Crystallographic data for other alkaline metal salts of thorium phosphates have been also reported.³ These compounds are monoclinic, space groups C2/c or Cc, $Z = 4$. Similar values of lattice constants and the same space group extinctions between all alkaline metal salts of thorium phosphates suggest close structural relationship among them. The crystal structure of $\text{KTh}_2(\text{PO}_4)_3$ has been solved in centrosymmetric space group C2/c and published recently.⁴ $\text{NaTh}_2(\text{PO}_4)_3$ ⁵ and $\text{NaU}_2(\text{PO}_4)_3$ ⁶ are piezoelectric, ferroelectric and therefore noncentrosymmetric, space group Cc.

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Experimental Section

To obtain single crystals of $\text{NaTh}_2(\text{PO}_4)_3$ previously described preparations were followed.^{2,3,4}

The crystals are monoclinic, space group Cc (No 9), with unit cell dimensions:

$$\begin{aligned} a &= 17.37 \pm .02 \text{ \AA} \\ b &= 6.81 \pm .01 \\ c &= 8.13 \pm .01 \\ \beta &= 101^\circ 3' \pm 10' \\ D_m &= 5.41; D_x = 5.435 \text{ g/cm}^3 \\ Z &= 4 \end{aligned}$$

The unit cell parameters were determined from oscillation and Weissenberg photographs with $\text{CuK}\alpha$ radiation. The density (D_m) was determined by a picnometer with decalin as liquid.

The three-dimensional intensity data were collected on multiple equi-inclination Weissenberg photographs using $\text{CuK}\alpha$ radiation. Three crystals were made spherical by grinding for absorption corrections ($\mu = 1133.4 \text{ cm}^{-1}$). On such ground specimens (with the radius of sphere $r = 0.081 \text{ mm}$, $\mu r = 9.2$; $r = 0.082 \text{ mm}$, $\mu r = 9.3$ and $r = 0.065 \text{ mm}$, $\mu r = 7.4$) following zones of reflections were collected: 0kl; h0l to h3l and hk0 to hk4. The relative intensities were determined from the optical densities of each spot by means of a microdensitometer. The corrections for absorption, polarization and Lorentz factors were made in the usual way and the structure amplitudes derived. The scale factors among the various hkl levels of data were improved in the course of structure determination. At the later stages of refinement the observed structure factors were transformed into absolute scale and the mean values were calculated for reflections whose intensities were determined several times. Weights were assigned using the method described by Wiesner and Lingafelter.⁷

Solution and Refinement of the Structure. The initial atomic coordinates, corresponding to those found in $\text{KTh}_2(\text{PO}_4)_3$,⁴ were assumed. The least squares refinement cycles were continued in C2/c space group and resulted in a high isotropic temperature factor

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(6) M. Topić and B. Prodić, *J. Appl. Cryst.*, **2**, 230 (1969).

(7) J. R. Wiesner and E. C. Lingafelter, *Inorg. Chem.*, **5**, 1170 (1966).

for the sodium atom. In the three-dimensional Fourier map the sodium atom peak was split, and suggested that the sodium atoms were disordered. The correct space group for $\text{NaTh}_2(\text{PO}_4)_3$ must be Cc because the crystals are piezoelectric.² When a non centric model was refined by least squares in the space group Cc with all the atoms in general positions, some of the isotropic temperature factors were negative and the coordinates of phosphorus and oxygen atoms were shifted so that P–O bond distances varied from 1.35 to 1.65 Å. Also some oxygen-oxygen distances between different PO_4 groupings were also too short, the shortest one was 2.48 Å. Since these irregularities in distances between the atoms could not have a physico-chemical explanation, a further refinement of the structure was attempted in the centric space group C2/c with the sodium atoms disordered about the twofold axis. The atomic coordinates and temperature factors obtained after several full-matrix least squares refinement cycles are listed in Table I and the observed and calculated structure factors in Table II. The atomic scattering factors used were from International Tables for X-ray Crystallography.⁸ Both, the real and imaginary, anomalous dispersion corrections were applied in the atomic scattering factor of Th atom. Isotropic thermal parameters for Th and P atoms, listed in Table I, are those

lengths and angles^{11,12} were carried out on the IBM 360-67 computer at the Washington State University Computing Center^{9,11} and on the CAE 90-40 computer at the Institute of Mathematics Computing Center (University of Zagreb).^{10,12} Table III gives the interatomic distances and angles obtained from the atomic coordinates in Table I. At this time it was found that the crystals of $\text{NaTh}_2(\text{PO}_4)_3$ possess ferroelectric properties.⁵ The samples investigated were thought to be polydomain crystals, and the observed structure an average. Several attempts were made to distinguish the positions of the sodium atoms in the domains assuming that the samples investigated were two-domain crystals and that all atoms except sodium met the symmetry requirements of C2/c. None of these models was significantly better on Hamilton's test¹³ than the disordered model.

Description and Discussion of the Structure

The main features of the structure of $\text{KTh}_2(\text{PO}_4)_3$ are repeated in $\text{NaTh}_2(\text{PO}_4)_3$. All PO_4 tetrahedra are discrete groups with no oxygen atom common to two phosphorus atoms, as is obvious from the formula of compound. Each oxygen atom, in addition to being bonded to a phosphorus atom, is also attached to the thorium atom. The packing of thorium atoms and

Table I. Fractional Coordinates for $\text{NaTh}_2(\text{PO}_4)_3$. Standard errors are given in parentheses

	x	y	z	B(Å ²)
Th(1)	0.1530(1)	0.0924(1)	0.0365(1)	0.59(3)
Na(2)	0.0184(31)	0.3883(83)	0.3482(64)	7.5 (1.4)
P(3)	0.3120(4)	0.0828(10)	0.3130(8)	0.7 (1)
P(4)	0	-0.1041(14)	1/4	1.1 (2)
O(5)	0.0752(11)	0.0297(31)	0.2530(26)	1.5 (4)
O(6)	-0.0240(13)	-0.2231(38)	0.0889(30)	2.2 (4)
O(7)	0.2270(10)	0.0517(26)	0.3310(21)	0.6 (2)
O(8)	0.3619(10)	-0.1017(28)	0.3392(23)	0.9 (3)
O(9)	0.2998(9)	0.1502(26)	0.1249(20)	0.5 (3)
O(10)	0.3486(11)	0.2644(34)	0.4127(26)	1.7 (4)

The anisotropic thermal parameters of Th and P atoms are of the form $T = \exp(-\beta_{11}h^2 + \beta_{22}k^2 + \beta_{33}l^2 + 2\beta_{12}hk + 2\beta_{13}hl + 2\beta_{23}kl)$. Each thermal parameter is multiplied by 10^4 . β_{12} and β_{23} for P(4) are required to be zero by symmetry.

	β_{11}	β_{22}	β_{33}	β_{12}	β_{13}	β_{23}
Th(1)	5(1)	26(2)	32(1)	0(0)	7(1)	1(1)
P(3)	10(2)	30(12)	22(8)	-7(4)	7(3)	3(9)
P(4)	7(3)	37(18)	80(15)	0	8(5)	0
$R_1 = \frac{\sum F_o - F_c }{\sum F_o } / \frac{\sum F_o }{\sum F_c } = 0.071$ (including unobserved reflections)						
$R_2 = \frac{\sum F_o - F_c }{\sum F_o } / \frac{\sum F_o }{\sum F_c } = 0.062$ (omitting unobserved reflections)						

obtained at $R_2=0.067$ (omitting unobserved reflections) before anisotropic temperature factors were adopted. Anisotropic temperature factors for Th and P were refined in the final stages of refinement. The refinement cycles^{9,10} and the calculation of bond

(8) International Tables for X-ray Crystallography, Vol. III, Birmingham: Kynoch Press (1962).

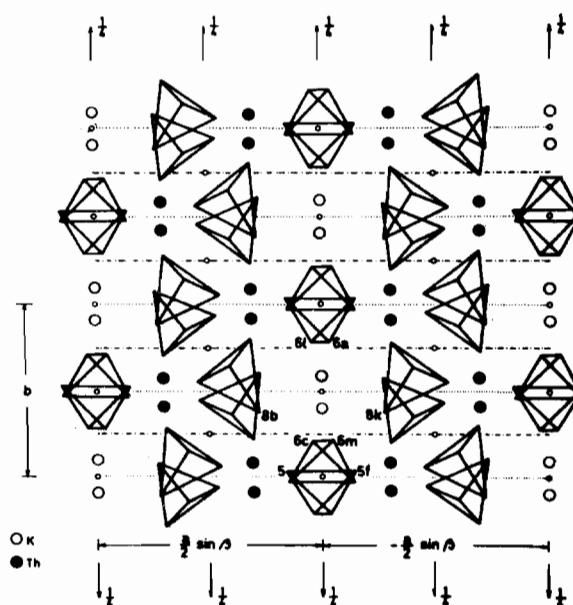


Figure 1. The structure of $\text{KTh}_2(\text{PO}_4)_3$ and symmetry elements projected down the c-axis. Oxygen atoms from the phosphate groups are connected by full lines. Oxygen atoms O(6c) and O(6m) belong to two different phosphate groups related by symmetry operation of twofold axis.

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(11) D. Anderson, A local version of the W. R. Busing, K. O. Martin and H. A. Levy, ORFFE, A Fortran crystallographic function and error program. U. S. Atomic Energy Commission Report ORNL-TM 306 (1964).

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(13) W. C. Hamilton, *Acta Cryst.*, 18, 502 (1965).

Table II. Observed, FO, and calculated, FC, structure factors

H	K	L	10FO	10FC	H	K	L	10FO	10FC	H	K	L	10FO	10FC	H	K	L	10FO	10FC	H	K	L	10FO	10FC	
0	0	2	3665	3837	2	2	1	2723	-2810	4	0	-8	1220	-1217	6	0	-6	734	-812	8	0	6	3457	-3349	
0	0	4	2458	2551	2	2	-1	6100	-6790	4	0	-10	493	414	6	0	8	618	606	8	0	-6	4014	3835	
0	0	6	948	932	2	2	2	1774	-1724	4	2	0	1540	-1616	6	0	-8	2942	-2799	8	0	8	2389	-2528	
0	0	8	1155	-1077	2	2	-2	0	233	4	2	1	3654	3884	6	0	-10	2321	-2543	8	0	-8	3770	3702	
0	0	10	1355	-1664	2	2	3	2408	-2625	4	2	-1	2190	2100	6	2	0	1290	1278	8	0	-10	1142	1364	
0	2	0	3266	3009	2	2	-3	4564	-704	4	2	2	1123	-992	6	2	1	1599	1640	8	2	0	0	-95	
0	2	1	1732	-1558	2	2	4	1148	-1136	4	2	-2	2381	-2183	6	2	-1	2402	2758	8	2	1	3491	-3453	
0	2	2	789	716	2	2	-4	1084	1039	4	2	3	3889	4091	6	2	2	1792	1561	8	2	-1	5668	-6000	
0	2	3	3024	-3075	2	2	5	912	-860	4	2	-3	48	-132	6	2	-2	1516	1491	8	2	2	619	-490	
0	2	4	1736	1785	2	2	-5	2885	-3024	4	2	4	391	-405	6	2	3	538	-559	8	2	-2	1662	1548	
0	2	5	3123	-3357	2	2	6	2119	-1971	4	2	-4	1539	-1462	6	2	-3	4956	5051	8	2	3	3444	-3595	
0	2	6	892	851	2	2	-6	1080	1118	4	2	5	4510	4740	6	2	4	1456	1476	8	2	-3	2549	-2463	
0	2	7	3624	-3747	2	2	7	1694	1574	4	2	-5	1769	-1743	6	2	-4	831	650	8	2	4	1071	-1028	
0	2	8	548	-480	2	2	-7	1007	-891	4	2	6	899	934	6	2	5	2270	-2220	8	2	-4	1357	1248	
0	2	9	1939	-1986	2	2	8	806	-757	4	2	-6	1506	-1402	6	2	-5	3400	3408	8	2	5	1917	-1893	
0	2	10	670	-769	2	2	-8	1945	1809	4	2	7	1750	1675	6	2	6	1198	1193	8	2	-5	820	-787	
0	4	0	3837	-3903	2	2	9	1776	1826	4	2	-7	2823	-2685	6	2	-6	0	-145	8	2	6	1312	-1228	
0	4	1	1575	-1471	2	2	-9	506	499	4	2	8	1010	961	6	2	7	2814	-2721	8	2	-6	1936	1741	
0	4	2	2531	-2366	2	2	-10	817	785	4	2	-8	871	-714	6	2	-7	3470	3514	8	2	7	382	-295	
0	4	3	2067	-1968	2	4	0	1514	1386	4	2	9	766	801	6	2	8	430	334	8	2	-7	591	504	
0	4	4	1809	-1703	2	4	1	2603	-2629	4	2	-9	2832	-2954	6	2	-8	1295	-1336	8	2	8	972	-961	
0	4	5	2533	-2536	2	4	-1	4107	-4236	4	2	-10	211	-6	6	2	9	1414	-1795	8	2	-8	1667	1561	
0	4	6	1017	-895	2	4	2	2408	2499	4	4	0	2329	2251	6	2	-9	1370	1371	8	2	-9	1679	1731	
0	4	7	2840	-2829	2	4	-2	0	-327	4	4	1	3052	3044	6	2	-10	947	-980	8	4	0	483	-505	
0	4	8	537	528	2	4	3	1998	-1820	4	4	-1	1961	2036	6	4	0	3296	-3370	8	4	1	2417	-2526	
0	4	9	1385	-1268	2	4	-3	2912	-2590	4	4	2	1348	1504	6	4	1	1346	952	8	4	-1	3669	-3606	
0	6	0	4557	-4669	2	4	4	2159	2024	4	4	-2	3156	3296	6	4	-1	1910	2077	8	4	2	1097	1098	
0	6	1	0	-163	2	4	-4	1449	-1420	4	4	3	2515	2522	6	4	2	2277	-2269	8	4	-2	1520	-1545	
0	6	2	2889	-2816	2	6	0	1319	1434	4	4	-3	0	-166	6	4	-2	1348	-1042	8	4	3	2794	-2812	
0	6	3	1368	1312	2	6	1	772	918	4	4	4	0	82	6	4	3	0	-387	8	4	-3	1635	-1496	
0	6	4	2124	-2052	2	6	-1	1602	1420	4	4	-4	2341	2111	6	4	-3	3908	3636	8	4	4	1443	1324	
0	6	5	1059	980	2	6	2	3190	2948	4	6	0	2791	2741	6	4	4	2807	-2629	8	4	-4	2951	-2891	
0	6	6	834	-801	2	6	-2	1289	-1076	4	6	1	1304	-1194	6	4	-4	1351	-749	8	6	0	692	-574	
0	6	7	904	836	2	6	3	993	927	4	6	-1	660	-490	6	6	0	3749	-3783	8	6	1	831	918	
0	8	0	241	-254	2	6	-3	959	1051	4	6	2	989	1016	6	6	1	0	-356	8	6	-1	1367	1268	
0	8	1	569	475	2	6	4	2976	2772	4	6	-2	4430	4307	6	6	-1	1196	-961	8	6	2	961	906	
0	8	2	230	-85	2	6	-4	1670	-1508	4	6	3	1096	-1041	6	6	2	2963	-2871	8	6	-2	2233	-2144	
0	8	3	1599	1778	2	8	0	360	385	4	6	-3	289	-214	6	6	-2	1743	-1749	8	6	3	1255	1125	
0	8	4	257	-106	2	8	1	2346	2323	4	6	4	0	-52	6	6	3	255	220	8	6	-3	1015	902	
1	1	0	3351	3302	2	8	-1	3437	3470	4	6	-4	2773	2679	6	6	-3	1627	-1477	8	6	4	2005	1909	
1	1	1	2815	-2595	2	8	2	0	-108	4	8	0	239	231	6	6	4	3037	-3152	8	6	-4	3015	-2858	
1	1	1	-2457	-2820	2	8	-2	545	-282	4	8	1	2661	-2674	6	6	-4	0	-326	8	8	0	489	234	
1	1	2	1413	1272	2	8	3	1311	1581	4	8	-1	1399	-1418	6	8	0	208	-97	8	8	1	2167	2350	
1	1	2	-2902	3266	2	8	-3	2028	2333	4	8	2	0	-52	6	8	1	759	-506	8	8	-1	2578	2978	
1	1	3	2623	-2738	3	1	0	3567	-3723	4	8	-2	0	217	6	8	-1	1718	-1779	8	8	2	0	-118	
1	1	3	-1206	-1365	3	1	1	0	-427	4	8	3	2265	-2360	6	8	2	0	-81	8	8	-2	318	-316	
1	1	4	1444	-1561	3	1	-1	1770	-1453	4	8	-3	294	236	6	8	-2	387	-249	9	1	0	2551	-2581	
1	1	4	-6250	6177	3	1	2	5441	-5840	5	1	0	0	-219	6	8	-3	3037	-3032	9	1	1	1510	-1439	
1	1	5	1612	-1556	3	1	-2	3761	-4486	5	1	1	3639	3570	7	1	0	3733	-3897	9	1	-1	2523	-2407	
1	1	5	-5	405	311	3	1	3	1126	1112	5	1	-1	1784	1806	7	1	1	1871	-1876	9	1	2	4056	-4069
1	1	6	2538	-2580	3	1	-3	1834	-1989	5	1	2	2359	2479	7	1	-1	736	-835	9	1	-2	1946	-1874	
1	1	6	2117	2145	3	1	4	2589	-2722	5	1	-2	1351	-1196	7	1	2	2781	2818	9	1	3	0	-218	
1	1	7	1588	-1619	3	1	-4	1842	-1966	5	1	3	1868	1852	7	1	-2	4960	4881	9	1	-3	2373	-2286	
1	1	7	-1121	1204	3	1	5	1904	1909	5	1	-3	2256	2236	7	1	3	2574	-2510	9	1	4	2970	-2918	
1	1	8	2178	-2043	3	1	-5	2858	-3128	5	1	4	3586	3618	7	1	-3	345	-457	9	1	-4	0	236	
1	1	8	-2018	1963	3	1	6	1569	-1531	5	1	-4	3132	-3290	7	1	4	666	580	9	1	5	647	587	
1	1	9	494	-432	3	1	-6	0	566	5	1	5	1017	1043	7	1	-4	3953	4100	9	1	-5	2949	-2909	
1	1	9	-1478	1464	3	1	7	1790	1780	5	1	-5	1440	1603	7	1	5	1561	-1338	9	1	6	2525	-2439	
1	1	10	2559	-2642	3	1	-7	1753	-1807	5	1	6	2984	2837	7	1	-5	1399	1449	9	1	6	2474	2514	
1	1	10	-10	571	742	3	3	8	0	-207	5	3	-6	1091	945	7	3	6	233	249	9	3	0	712	648
1	1	11	0	-509	296	3	3	7	187	-206	5	3	-6	0	-109	7	3	8	0	125	11	7	1	2423	-2598
1	1																								

Table II. (Continued)

H	K	L	10FO	10FC	H	K	L	10FO	10FC	H	K	L	10FO	10FC	H	K	L	10FO	10FC	H	K	L	10FO	10FC					
12	2	-8	1178	-1185	13	3	-4	0	-151	14	4	-1	1827	-1725	15	5	-1	582	-719	17	1	-5	807	-782	19	1	1	707	675
12	2	-9	410	320	13	3	5	2249	-2106	14	4	2	425	-458	15	5	-2	829	-762	17	1	-6	2558	-2410	19	1	-1	1363	1277
12	4	0	1289	-1221	13	3	-5	3451	3106	14	4	-2	2286	-2176	15	5	3	474	-359	17	1	-7	1580	-1646	19	1	2	2196	2176
12	4	1	1778	1697	13	3	6	138	49	14	4	3	2460	-2647	15	5	-3	590	-548	17	1	-8	735	-698	19	1	-2	1164	1149
12	4	-1	2430	2545	13	3	-6	222	-79	14	4	-3	893	-866	15	5	-4	1720	-1793	17	3	0	442	321	19	1	3	278	-260
12	4	2	1895	-1881	13	3	-7	3783	3773	14	4	4	505	569	16	0	0	3079	-3119	17	3	1	2991	2817	19	1	-3	1804	1763
12	4	-2	0	-221	13	3	-8	0	63	14	4	-4	2614	-2663	16	0	2	4002	-3960	17	3	-1	1154	1084	19	1	-4	186	169
12	4	5	1304	1287	13	5	0	3510	-3764	14	6	0	1583	-1727	16	0	-2	3564	-3373	17	3	2	371	318	19	1	-5	1615	1642
12	4	-3	3775	3591	13	5	1	592	-527	14	6	1	624	731	16	0	4	1956	-1991	17	3	-2	773	744	19	1	-6	705	-718
12	4	4	2266	-2266	13	5	-1	322	402	14	6	-1	1015	961	16	0	-4	1202	-1113	17	3	3	2360	2459	19	3	0	598	-536
12	4	-4	677	642	13	5	2	2336	-2416	14	6	-2	2360	-2371	16	0	-6	785	773	17	3	-3	362	-235	19	3	1	1002	1137
12	6	0	1692	-1723	13	5	-2	2746	-2730	14	6	-3	310	-197	16	0	-8	1991	1807	17	3	4	0	54	19	3	-1	2209	2292
12	6	1	779	-1132	13	5	3	0	-144	14	6	-4	2601	-3242	16	2	0	1039	-994	17	3	-4	531	493	19	3	2	428	-357
12	6	-1	1090	-974	13	5	-3	343	267	15	1	0	1148	-1169	16	2	1	576	-309	17	3	-5	1541	-1587	19	3	-2	0	-93
12	6	2	2154	-2136	13	5	4	1447	-1832	15	1	1	1647	-1597	16	2	-1	1680	-1602	17	3	-6	400	473	19	3	-3	2835	2996
12	6	-2	382	-415	13	5	-4	2044	-2044	15	1	-1	2696	-2479	16	2	2	1745	-1553	17	3	-7	2544	-2619	19	3	-4	0	-48
12	6	3	386	-1119	13	7	0	1780	-1964	15	1	2	2491	-2171	16	2	-2	1216	-1029	17	5	0	1770	2050	19	3	-5	2316	2671
12	6	-3	1220	-1171	13	7	-1	697	-543	15	1	-2	537	683	16	2	3	840	786	17	5	-1	369	290	20	0	0	2724	2971
12	6	-4	1453	1516	13	7	-2	1400	-1290	15	1	3	1395	-1003	16	2	-3	3076	-2927	17	5	-2	3019	3253	20	0	2	1800	2081
13	1	0	4582	4663	14	0	0	2931	3070	15	1	-3	1843	-1685	16	2	4	901	-825	17	5	-3	109	-70	20	0	-2	2859	2890
13	1	1	408	-555	14	0	2	1068	881	15	1	4	1964	-1816	16	2	-4	524	-542	17	5	-4	2147	2374	20	0	-4	3185	3188
13	1	-1	1216	972	14	0	-2	3211	3220	15	1	-4	2285	2220	16	2	5	1849	1791	18	0	0	397	-310	20	0	-6	1350	1305
13	1	2	2361	2323	14	0	4	1050	-1094	15	1	5	386	-378	16	2	-5	3623	-3657	18	0	2	1446	1375	20	2	0	946	912
13	1	-2	2522	2447	14	0	-4	4949	4899	15	1	-5	1725	-1549	16	2	-6	0	-60	18	0	-2	1494	-1427	20	2	1	1160	-1305
13	1	3	1527	-1384	14	0	6	1772	-1635	15	1	6	2585	-2566	16	2	-7	2141	-2302	18	0	-4	2228	2557	20	2	-1	553	-523
13	1	-3	1400	1403	14	0	-6	2642	2612	15	1	-6	2456	2400	16	2	-8	705	614	18	0	-4	2766	-2735	20	2	2	827	888
13	1	4	2034	2025	14	0	-8	2197	2240	15	1	-7	615	-376	16	4	0	1595	1619	18	0	-3	3348	-3365	20	2	-2	1203	1195
13	1	-4	2207	2107	14	2	0	1217	1072	15	1	-8	2818	2869	16	4	-1	1235	-1273	18	2	0	0	85	20	2	-3	984	1037
13	1	5	1143	-1077	14	2	1	2939	-2741	15	3	0	0	342	16	4	-2	2112	1916	18	2	1	2955	2972	20	2	-4	1252	1291
13	1	-5	1899	1855	14	2	-1	2925	-2820	15	3	1	2736	-2799	16	4	3	815	898	18	2	-1	2622	2610	20	2	-5	2012	1969
13	1	6	781	709	14	2	2	805	452	15	3	-1	4355	-4246	16	4	-3	2166	-2168	18	2	2	520	527	21	1	0	176	314
13	1	-6	442	429	14	2	-2	1096	1022	15	3	2	540	424	16	4	-4	872	801	18	2	-2	388	-341	21	1	1	1401	-1338
13	1	7	1342	-1412	14	2	3	3446	-3489	15	3	-2	0	-120	16	6	-1	607	533	18	2	3	1697	1862	21	1	-1	1710	-1688
13	1	-7	2471	2430	14	2	-3	696	-607	15	3	3	1625	-1563	16	6	-2	2113	2104	18	2	-3	3009	3008	21	1	-2	1880	2039
13	1	-8	983	-887	14	2	4	484	-400	15	3	-3	3165	-2954	16	6	-3	407	894	18	2	-4	1204	-1158	21	1	-3	706	-746
13	1	-9	1169	1219	14	2	-4	2236	2061	15	3	4	274	295	17	1	0	2138	-2154	18	2	-5	1064	1072	21	1	-4	1806	2242
13	3	0	766	-679	14	2	5	1950	-1996	15	3	-4	412	-531	17	1	1	1522	1342	18	2	-6	1196	-1251	21	1	-5	517	-420
13	3	1	834	-823	14	2	-5	1462	1311	15	3	5	473	-507	17	1	-1	690	640	18	2	-7	148	142	21	3	-2	0	-285
13	3	-1	2094	2006	14	2	6	598	-688	15	3	-5	2662	-2658	17	1	2	1280	-1220	18	4	0	0	56	22	0	-2	1088	-995
13	3	2	906	-757	14	2	-6	1160	1073	15	3	-6	529	-546	17	1	-2	3514	-3397	18	4	-1	1993	2069					
13	3	-2	547	-589	14	2	-7	1742	1706	15	3	-7	587	-622	17	1	3	1331	1256	18	4	-2	1085	698					
13	3	3	2360	-2280	14	2	-8	981	880	15	3	-8	427	-421	17	1	-3	0	287	18	4	-3	2031	2107					
13	3	-3	2534	2491	14	4	0	1466	-1619	15	5	0	949	928	17	1	4	228	-279	18	4	-4	1290	1402					
13	3	4	255	-328	14	4	1	1821	-1785	15	5	1	605	-668	17	1	-4	2682	-2658	19	1	0	2608	2567					

Figure 2. A view down the polyhedron of oxygen atoms around the alkali metal atom (which is omitted in the figure).

In the structure of $\text{KTh}_2(\text{PO}_4)_3$ potassium atom is located on the twofold axis close to the centre of the cavity.

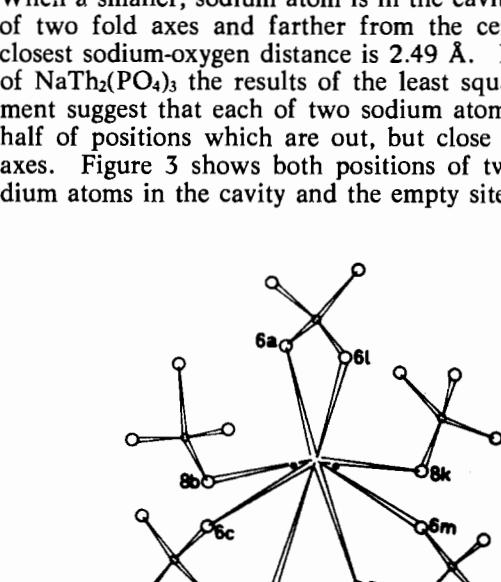


Figure 3. The positions of two sodium atoms (black circles) in the structure of $\text{NaTh}_2(\text{PO}_4)_3$ in comparison with the empty site where, in $\text{KTh}_2(\text{PO}_4)_3$, potassium atom is located. The empty site of potassium atom and coordinated oxygen atoms from phosphates groups are connected by full lines.

Table III. Interatomic distances and angles in $\text{NaTh}_2(\text{PO}_4)_3$. Standard errors are in parentheses

A	B	C	Distances (\AA)		Angles (degrees) A-B-C
			A-B	A-C	
Within Th-polyhedron					
O(6c)	Th(1)	O(5e)	2.44(2)	2.42(3)*	57.5(7)
O(7)	Th(1)	O(9)	2.51(2)	2.38(2)*	56.3(4)
O(5e)	Th(1)	O(7e)	2.58(2)	2.60(3)	61.5(6)
O(5)	Th(1)	O(7)	2.45(2)	2.60(3)	63.1(6)
O(9g)	Th(1)	O(9)	2.42(1)	2.76(2)	67.4(4)
O(9)	Th(1)	O(7e)	2.54(1)	2.84(2)	68.5(4)
O(7e)	Th(1)	O(9g)	2.50(2)	2.81(2)	69.6(5)
O(8b)	Th(1)	O(9g)	2.35(2)	2.76(1)	70.6(4)
O(10h)	Th(1)	O(5)	2.27(1)	2.74(2)	70.8(6)
O(5)	Th(1)	O(6c)		3.05(3)	77.0(7)
O(8b)	Th(1)	O(6c)		2.82(3)	71.9(7)
O(7)	Th(1)	O(10h)		2.91(2)	74.9(4)
O(10h)	Th(1)	O(7e)		2.95(2)	76.2(5)
O(8b)	Th(1)	O(5)		2.89(3)	74.0(6)
O(7)	Th(1)	O(8b)		3.01(2)	76.5(5)
O(10h)	Th(1)	O(5e)		3.12(2)	79.8(5)
O(8b)	Th(1)	O(9)		3.34(2)	85.9(6)
O(5e)	Th(1)	O(9g)		3.40(3)	85.6(5)
O(6c)	Th(1)	O(9g)		3.25(3)	83.6(7)
O(10h)	Th(1)	O(9)		3.65(2)	98.5(6)
O(5)	Th(1)	O(5e)		4.09(1)	108.6(7)
O(7)	Th(1)	O(7e)		4.13(1)	111.1(6)
O(10h)	Th(1)	O(6c)		3.94(3)	113.0(8)
O(6c)	Th(1)	O(7e)		4.15(3)	114.4(7)
O(7)	Th(1)	O(9g)		4.17(2)	115.6(5)
O(5)	Th(1)	O(9)		4.31(2)	119.0(5)
O(8b)	Th(1)	O(5e)		4.39(2)	126.0(5)
O(5e)	Th(1)	O(9)		4.62(2)	128.9(5)
O(7)	Th(1)	O(6c)		4.56(3)	134.4(7)
O(8b)	Th(1)	O(7e)		4.53(2)	138.5(6)
O(10h)	Th(1)	O(8b)		4.37(2)	141.9(3)
O(5)	Th(1)	O(9g)		4.63(2)	143.5(6)
O(10h)	Th(1)	O(9g)		4.49(2)	145.7(5)
O(5)	Th(1)	O(7e)		4.74(3)	146.8(6)
O(6c)	Th(1)	O(9)		4.80(3)	148.3(7)
O(7)	Th(1)	O(5e)		4.96(3)	154.7(6)

Within the phosphate groups

O(6)	P(4)	O(5f)	1.53(3)	2.42(3)	102(1)
O(6f)	P(4)	O(5)			
O(5)	P(4)	O(6)			
O(5f)	P(4)	O(6f)	1.59(2)	2.61(3)	114(1)
O(5)	P(4)	O(5f)		2.60(4)	110(1)
O(6)	P(4)	O(6f)		2.59(5)	116(2)
O(7)	P(3)	O(9)	1.53(2)	2.38(2)	100(1)
O(9)	P(3)	O(10)	1.57(1)	2.46(1)	104(1)
O(8)	P(3)	O(7)	1.52(2)	2.55(2)	114(1)
O(10)	P(3)	O(7)	1.55(1)	2.54(2)	112(1)
O(8)	P(3)	O(9)		2.53(2)	110(1)
O(8)	P(3)	O(10)		2.58(2)	115(1)

* Also within the phosphate group.

mean values from the least squares refinement carried out in the space group C2/c. The $[\text{Th}_2(\text{PO}_4)_3]$ groupings in $\text{NaTh}_2(\text{PO}_4)_3$ are not significantly different from those in $\text{KTh}_2(\text{PO}_4)_3$. The interatomic distances and angles in $\text{NaTh}_2(\text{PO}_4)_3$ are represented in Table III. Thorium-oxygen distances are from 2.27 to 2.58 \AA . Phosphorus-oxygen distances within the phosphate groups vary from 1.52 to 1.59 \AA . From the results obtained in the idealized structure it is obvious that the bond lengths and angles in PO_4 tetrahedra deviate to some small extent from the results obtained and discussed before.⁴

Table III. (Continued)

Sodium-oxygen distances $\leq 4.02 \text{\AA}$		
Standard errors are 0.06 \AA		
Na(2)	O(6m)	2.49
Na(2f)	O(6c)	
Na(2)	O(8k)	2.71
Na(2f)	O(8b)	
Na(2)	O(6l)	2.69
Na(2f)	O(6a)	
Na(2)	O(5)	2.80
Na(2f)	O(5f)	
Na(2)	O(8b)	2.80
Na(2f)	O(8k)	
Na(2)	O(10j)	2.91
Na(2f)	O(10d)	
Na(2)	O(5f)	2.96
Na(2f)	O(5)	
Na(2)	O(8j)	3.30
Na(2f)	O(8d)	
Na(2)	O(6a)	3.38
Na(2f)	O(6l)	
Na(2)	O(9b)	3.60
Na(2f)	O(9k)	
Na(2)	O(6c)	3.75
Na(2f)	O(6m)	
Na(2)	O(10k)	4.02
Na(2f)	O(10b)	

Numbering of the Atoms

Small letters indicate symmetry transformations or cell translations:

a: $x, y+1, z$ h: $\frac{1}{2}-x, \frac{1}{2}+y-1, \frac{1}{2}-z$
 b: $\frac{1}{2}-x, \frac{1}{2}+y, \frac{1}{2}-z$ i: $\frac{1}{2}+x-1, \frac{1}{2}+y-1, z$
 c: $-x, -y, -z$ j: $\frac{1}{2}-x, \frac{1}{2}-y, 1-z$
 d: $\frac{1}{2}+x-1, \frac{1}{2}-y, \frac{1}{2}+z-1$ k: $\frac{1}{2}+x-1, \frac{1}{2}+y, z$
 e: $x, -y, \frac{1}{2}+z-1$ l: $-x, y+1, \frac{1}{2}-z$
 f: $-x, y, \frac{1}{2}-z$ m: $x, -y, \frac{1}{2}+z$
 g: $\frac{1}{2}-x, \frac{1}{2}-y, -z$

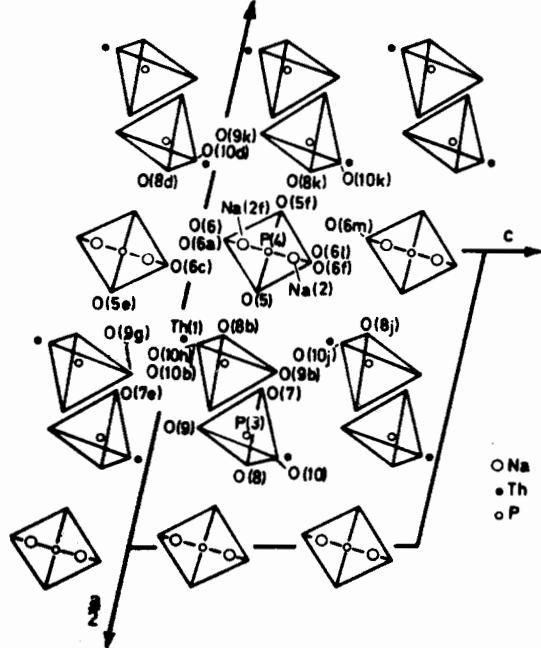


Figure 4. Projection of the structure of $\text{NaTh}_2(\text{PO}_4)_3$ on a plane normal to b-axis. Oxygen atoms from the phosphate groups are connected by full lines. Only those atoms are numbered which are mentioned in the text or tables.

The present paper describes without any ambiguity the structural model of a hitherto unknown ferroelectric compound. Bond lengths and angles do not represent the absolute truth because they were obtained on the assumption that the shifts of atoms in domains do not have significant value. The standard deviations of bond lengths and angles, which were obtained from

the least squares refinement, are probably under estimated. More precise structural studies of a monodomain ferroelectric sample stabilized by an electric field are planned in order to give more accurate values for bond lengths and angles and to completely explain why the crystals of $\text{NaTh}_2(\text{PO}_4)_3$ exhibit ferroelectric properties.